

## Research Article

# Fabrication of $\text{CuO}_x$ -Pd Nanocatalyst Supported on a Glassy Carbon Electrode for Enhanced Formic Acid Electro-Oxidation

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Received 27 August 2018; Revised 17 October 2018; Accepted 13 November 2018; Published 2 December 2018

Guest Editor: Abdellatif Ait Lahcen

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Formic acid (FA) electro-oxidation (FAO) was investigated at a binary catalyst composed of palladium nanoparticles (PdNPs) and copper oxide nanowires ( $\text{CuO}_x$ NWs) and assembled onto a glassy carbon (GC) electrode. The deposition sequence of PdNPs and  $\text{CuO}_x$ NWs was properly adjusted in such a way that could improve the electrocatalytic activity and stability of the electrode toward FAO. Several techniques including cyclic voltammetry, chronoamperometry, field-emission scanning electron microscopy, energy dispersive X-ray spectroscopy, and X-ray diffraction were all combined to report the catalyst's activity and to evaluate its morphology, composition, and structure. The highest catalytic activity and stability were obtained at the  $\text{CuO}_x$ /Pd/GC electrode (with PdNPs directly deposited onto the GC electrode followed by  $\text{CuO}_x$ NWs with a surface coverage,  $\Gamma$ , of ca. 49%). Such enhancement was inferred from the increase in the peak current of direct FAO (by ca. 1.5 fold) which associated a favorable negative shift in its onset potential (by ca. 30 mV). The enhanced electrocatalytic activity and stability (decreasing the loss of active material by ca. 1.5-fold) of the  $\text{CuO}_x$ /Pd/GC electrode was believed originating both from facilitating the direct oxidation (decreasing the time needed to oxidize a complete monolayer of FA, increasing turnover frequency, by ca. 2.5-fold) and minimizing the poisoning impact (by ca. 71.5%) at the electrode surface during FAO.

## 1. Introduction

The growing petition for a clean fuel and the awareness of environmental problems have drawn the attention of the global community to substitute fossil fuels with other sources of clean energy [1, 2]. Of these, fuel cells were encouraging as they have been verified competent, consistent, sustainable, noiseless, long-lasting, and simply installed and moved; hence, they were excellent for transportation and suburban uses as well as portable electronics [3–7]. To date, the direct formic acid fuel cells (DFAFCs) have shown more preferred than hydrogen fuel cells (HFCs) and direct methanol fuel cells (DMFCs) in providing electricity for portable electronics [6–14]. In fact, the transportation and storage of formic acid (FA, the energy carrier in DFAFCs)

was much easier than  $\text{H}_2$ , and it could exhibit, moreover, a lower crossover than methanol through Nafion-based membranes. This permitted a higher practical energy density because of allowing the usage of thinner membranes and more concentrated fuel solutions [15, 16]. In addition, FA can be sustainably manufactured by  $\text{CO}_2$  electrochemical reduction utilizing excess clean electrical power sources such as solar farms and wind turbines [17].

With the revolution of nanotechnology, several nanostructures have been recommended for applications of multidiversity [18–21]. Platinum-based catalysts have been commonly used as for the electro-oxidation of formic acid (FAO) [22, 23]. On Pt-based materials, FAO proceeded in two pathways at the same time: the direct (favorable—in which FA is converted to  $\text{CO}_2$  at a low anodic potential)

[9, 15–17] and the indirect (unfavorable—in which FA is converted to poisoning CO that is then oxidized at higher potentials by platinum hydroxide species) [24]. Unluckily, CO adsorption in the low potential domain at the Pt surface deactivated its catalytic activity, which finally impeded the direct pathway of FAO [24–26]. Scheme 1 represents the dehydrogenation and dehydration pathways of FAO at Pt surfaces.

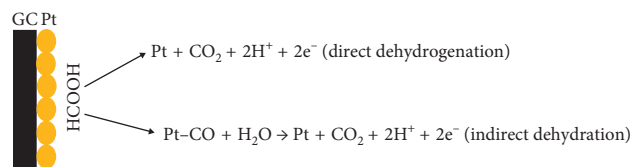
Compared with Pt, Pd can catalyze the FAO through a more facile direct path to produce CO<sub>2</sub> with less vulnerability to CO poisoning [27, 28]. However, Pd-based catalysts used to be subjected to a potential deactivation with time that originated typically from the adsorption of some CO-like poisoning species [29, 30]. The catalytic performance of Pd catalysts toward FAO was recently enhanced by doping with other metals and/or metal oxides that have the capacity to reduce the adsorption of these poisoning intermediates [31–34]. In the previous study [12], MnO<sub>x</sub> was able to enhance the electrocatalytic performance of Pd in such a way that minimized the adsorption of poisoning species depicted from calculating the long-term poisoning rate ( $\delta$ ). This modification reduced  $\delta$  to 0.11, 0.08, and 0.05%·s<sup>-1</sup> at the Pd/MnO<sub>x</sub>/GC, MnO<sub>x</sub>/Pd/GC, and Pd-MnO<sub>x</sub>/GC electrodes, respectively.

Doping platinum and palladium with nonprecious copper or copper oxide has not only exhibited outstanding improvement in the catalytic activities and stability of several potential electrochemical reactions [35–38], but also it offers a very effective way that could reduce the loading of noble metals. Hence, a motivation to report on the enhanced efficiency of a CuO<sub>x</sub>/Pd/GC nanocatalyst (in which PdNPs was directly electrodeposited onto a GC surface followed by CuO<sub>x</sub>NWs) toward FAO exists. The electrodeposition technique is utilized as it attracted particular attention due to the ease of preparation, suitability for special-shaped electrodes, and low-cost requirement. The influence of deposition sequence and loading level of CuO<sub>x</sub>NWs was examined to maximize the catalytic performance toward FAO.

## 2. Experimental

**2.1. Electrochemical Measurements.** After cleaning with conventional procedures, the electrode was mechanically polished with No. 2000 emery paper and then with aqueous slurries of successively finer alumina powder (down to 0.06  $\mu$ m) with the help of a polishing microcloth. Next, the polished electrode was rinsed thoroughly with distilled water, and glassy carbon (GC,  $d = 3.0$  mm) electrode was served as a working electrode. A spiral Pt wire and an Ag/AgCl/KCl (sat.) were used as counter and reference electrodes, respectively. All potentials in this investigation were measured in reference to Ag/AgCl/KCl (sat.).

The electrochemical measurements were performed at room temperature ( $25 \pm 1$  °C) in a two-compartment three-electrode glass cell. The measurements were performed using EG&G potentiostat (model 273A). Current densities were calculated on the basis of the real surface area of the working electrodes. The electrocatalytic activity of the



SCHEME 1: Dehydrogenation and dehydration pathways of FAO.

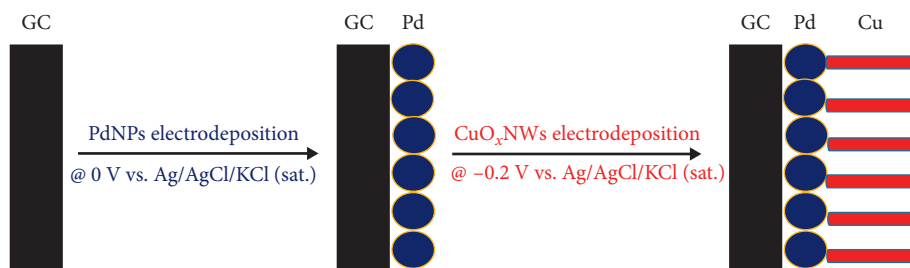
modified electrodes toward FAO was examined in an aqueous solution of 0.3 M FA solution (pH = 3.5, pH was adjusted by adding a proper amount of NaOH) at a scan rate of 100 mV·s<sup>-1</sup> (first cycle is recorded). At this pH, the polarization resistance will be reduced, thus the solution ionic conductivity was enhanced [39].

**2.2. Electrodeposition of PdNPs and CuO<sub>x</sub>NWs.** The electrodeposition of PdNPs was carried out in 0.1 M H<sub>2</sub>SO<sub>4</sub> solution containing 1.0 mM Pd(CH<sub>3</sub>COO)<sub>2</sub> at a constant potential electrolysis at 0 V for 300 s. The PdNPs loading was estimated from Faraday's law using the charge associated in the  $i$ - $t$  curve obtained during deposition to be ca. 18  $\mu$ g·cm<sup>-2</sup>. Alternatively, the electrode's modification with CuO<sub>x</sub>NWs was carried out in 0.1 M H<sub>2</sub>SO<sub>4</sub> solution containing 1.0 mM CuSO<sub>4</sub> at a constant potential electrolysis at -0.2 V for various durations [38]. These durations were set to control the amount of the electrodeposited Cu. The experimental protocol used for the surface modification of GC electrode with PdNPs and CuNWs is shown in Scheme 2.

**2.3. Materials Characterization.** A field-emission scanning electron microscope, FE-SEM, (QUANTA FEG 250) coupled with an energy dispersive X-ray spectrometer (EDS) was served to determine the morphology and composition of the investigated electrodes. An X-ray diffraction (XRD, PANalytical, X'Pert PRO) operated with Cu target ( $\lambda = 1.54$  Å) revealed the crystallographic structure of the modified catalysts. The inductively coupled plasma optical emission spectrometry, ICP-OES, (Perkin Elmer, Optima 8000) has been utilized to specify if there was a loss in the catalysts' materials (Pd&Cu) after the prolonged electrolysis and to determine the amounts lost.

## 3. Results and Discussion

**3.1. Electrochemical and Materials Characterization.** Cyclic voltammograms (CVs) for (a) Pd/GC, (b) CuO<sub>x</sub>/GC, (c) Pd/CuO<sub>x</sub>/GC, and (d) CuO<sub>x</sub>/Pd/GC electrodes measured in 0.5 M H<sub>2</sub>SO<sub>4</sub> are shown in Figure 1. Figure 1(a) depicts the characteristic response of Pd surfaces where its oxidation started between 0.6 and 1.2 V was combined with this oxide reduction at ca. 0.4 V [33]. Besides that, the peaks that appeared in the potential range from 0 to -0.2 V were assigned to the hydrogen adsorption/desorption (H<sub>ads/des</sub>). For the CuO<sub>x</sub>/GC electrode (Figure 1(b)), a main oxidation peak appeared at ca. 0.15 V coupled with a reduction peak after ca. 0 V corresponding to the CuO<sub>x</sub> formation and reduction [27, 29]. After the deposition of PdNPs onto the



SCHEME 2: The experimental protocol used for the surface modification of GC electrode with PdNPs and CuNWs.

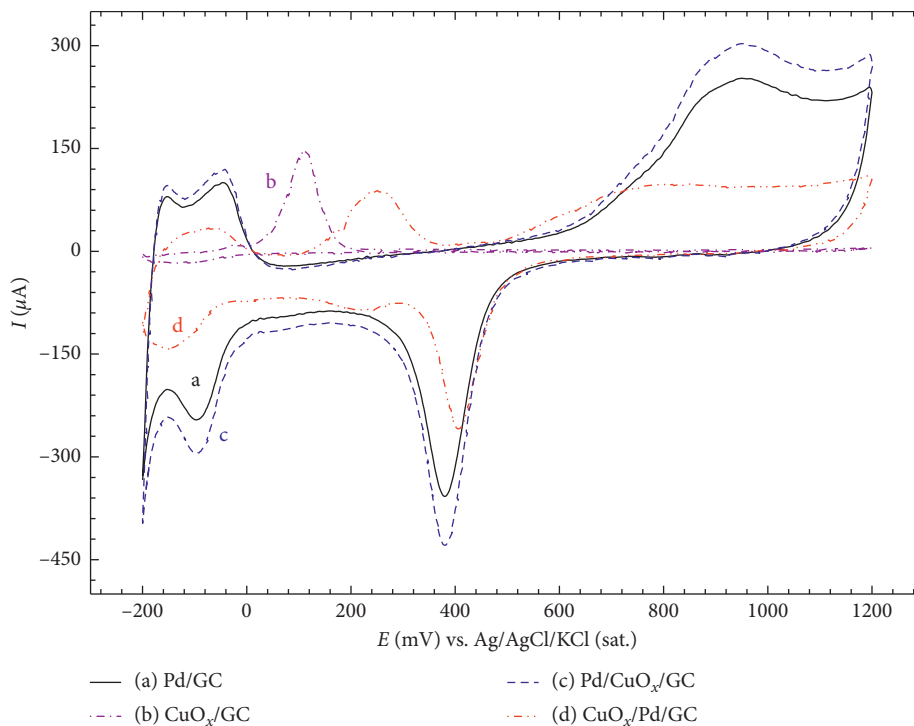


FIGURE 1: CVs obtained at (a) Pd/GC, (b) CuO<sub>x</sub>/GC, (c) Pd/CuO<sub>x</sub>/GC, and (d) CuO<sub>x</sub>/Pd/GC electrodes in 0.5 M H<sub>2</sub>SO<sub>4</sub>. Potential scan rate: 100 mVs<sup>-1</sup>.

CuO<sub>x</sub>/GC electrode (Pd/CuO<sub>x</sub>/GC electrode, Figure 1(c)), the intensity of the Pd oxide reduction peak increased coinciding with the increase in the intensity of the Pd oxide and the H<sub>ads/des</sub> peaks. This is expected as the loading of CuO<sub>x</sub>NWs on the GC surface can offer more active sites for the deposition of PdNPs; therefore, the Pd surface area could be seen increased. On the contrary, upon modifying the Pd/GC electrode with CuO<sub>x</sub>NWs (Figure 1(d)), two new features were noticed:

- (i) A significant decrease in the intensities of the Pd oxide reduction and the H<sub>ads/des</sub> peaks. This suggests a significant decrease in the PdNPs surface area as a result of CuO<sub>x</sub>NWs deposition.
- (ii) A redox couple at ca. 0.25 V corresponding to the Cu oxidation and reduction [27–30].

These new features of the CuO<sub>x</sub>/Pd/GC catalyst confirmed the surface exposure of both PdNPs and CuO<sub>x</sub>NWs to the electrolyte.

Figure 2 displays FE-SEM images of the Pd/GC (Figure 2(a)), CuO<sub>x</sub>/GC (Figure 2(b)), Pd/CuO<sub>x</sub>/GC (Figure 2(c)), and CuO<sub>x</sub>/Pd/GC (Figure 2(d)) electrodes. It illustrates that Pd was electrodeposited onto the GC electrode in a well-dispersed spherical lumps with an average particle size of ca. 75 nm (Figures 2(a), 2(c), and 2(d)), whereas CuO<sub>x</sub> has been electrodeposited as hair-like nanowires (Figures 2(b) and 2(d)). The disappearance of the CuO<sub>x</sub> morphology for the Pd/CuO<sub>x</sub>/GC electrode (Figure 2(c)), in contrast to the case for CuO<sub>x</sub>/Pd/GC (Figure 2(d)), infers that it has been covered during Pd deposition. This difference in morphology may affect the electrocatalytic parameters of FAO which will be displayed in the next section.

The EDS analysis of the Pd/GC, CuO<sub>x</sub>/GC, and CuO<sub>x</sub>/Pd/GC electrodes confirmed the deposition of the different ingredients in the catalyst and assisted in calculation of their relative ratios (Figures 3(a)–3(c)). Moreover, the crystal structure of the Pd/GC, CuO<sub>x</sub>/GC, and CuO<sub>x</sub>/Pd/GC electrodes was investigated utilizing the XRD technique.

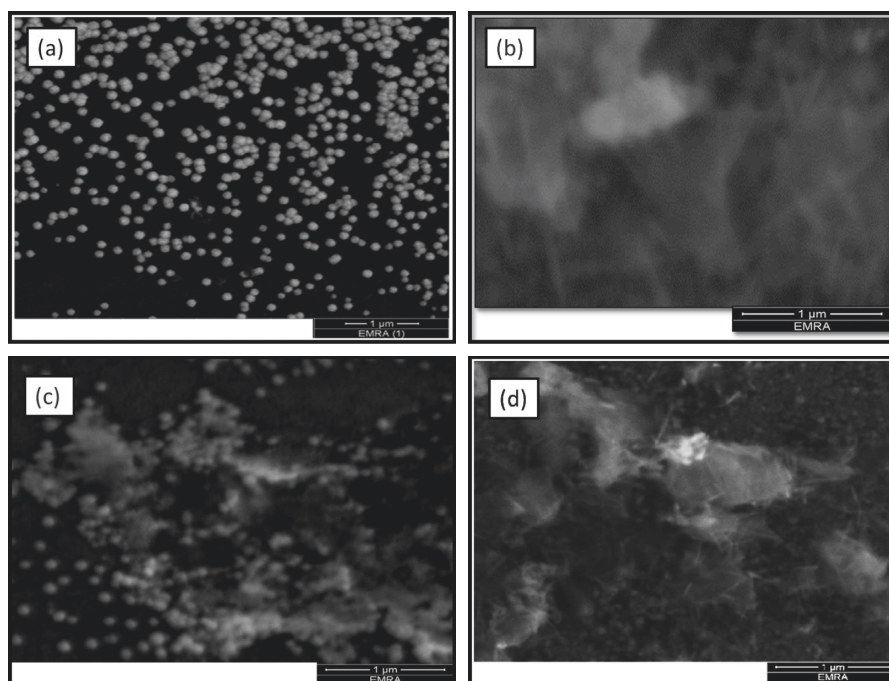


FIGURE 2: FE-SEM images of (a) Pd/GC, (b) CuO<sub>x</sub>/GC, (c) Pd/CuO<sub>x</sub>/GC, and (d) CuO<sub>x</sub>/Pd/GC electrodes. The electrodeposition conditions are listed in the experimental section.

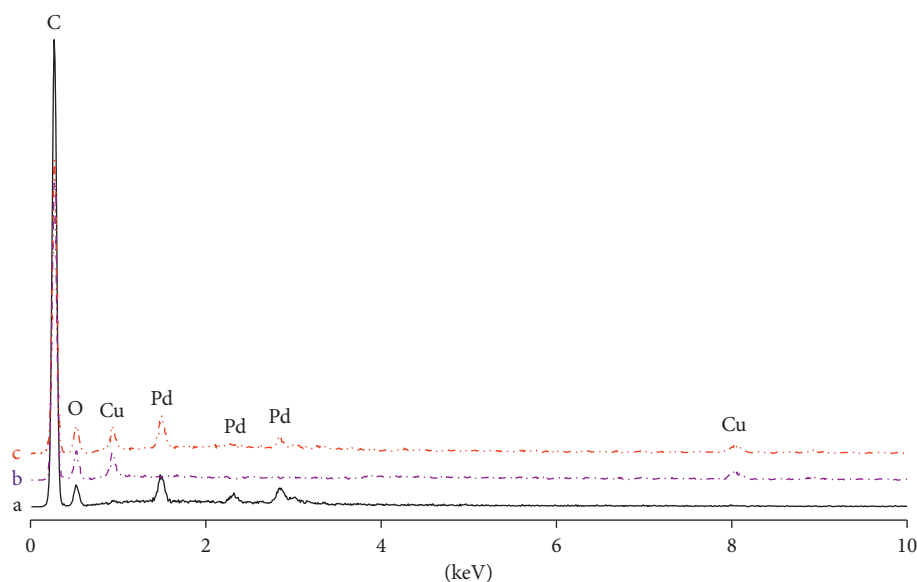


FIGURE 3: EDS analysis of (a) Pd/GC, (b) CuO<sub>x</sub>/GC, and (c) CuO<sub>x</sub>/Pd/GC electrodes. The electrodeposition conditions are listed in the experimental section.

In fact, the XRD analysis of the Pd/GC and CuO<sub>x</sub>/Pd/GC electrodes (Figures 4(a) and 4(c)) displayed several diffraction peaks at ca. 39.7°, 43.2°, 67.2°, and 78° corresponding, respectively, to the (1 1 1), (2 0 0), (2 2 0), and (3 1 1) planes of Pd face-centered cubic (fcc) lattice [40, 41]. However, in Figures 4(b) and 4(c), a very small diffraction peak appeared at ca. 43.3° corresponding to Cu (1 1 1) plane. There were also two broad diffraction peaks for C (0 0 2) and (1 0 0) planes between 21–28° and 42–46°, respectively [42]. The C (1 0 0)

peak was much higher and broader than those of Pd (200) and Cu (111) and that is why they appeared like a single peak.

### 3.2. Electrocatalytic Activity toward FAO

**3.2.1. Effect of Deposition Sequence.** Figure 5 shows CVs of FAO at the Pd/GC, CuO<sub>x</sub>/GC, Pd/CuO<sub>x</sub>/GC, and CuO<sub>x</sub>/Pd/GC electrodes in a 0.3 M aqueous solution of FA (pH = 3.5).



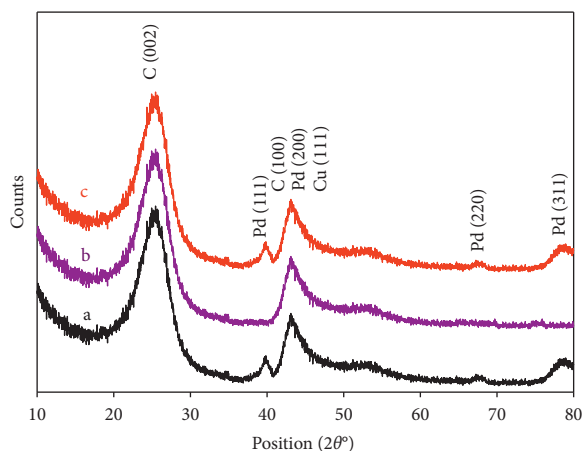


FIGURE 4: XRD patterns of (a) Pd/GC, (b)  $\text{CuO}_x$ /GC, and (c)  $\text{CuO}_x$ /Pd/GC electrodes. The electrodeposition conditions are listed in the experimental section.

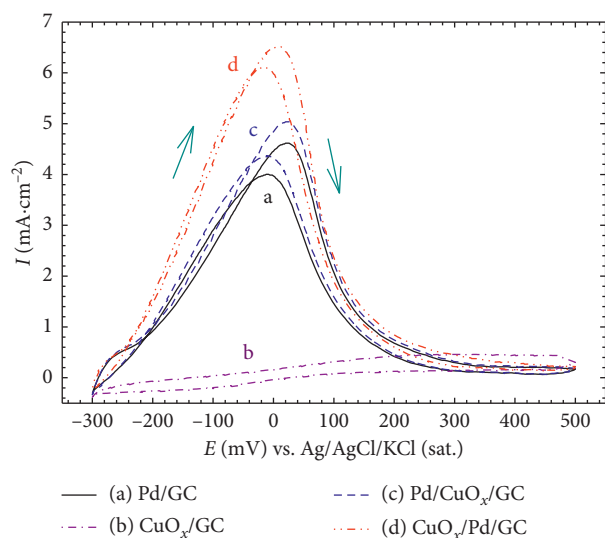


FIGURE 5: CVs obtained at (a) Pd/GC, (b)  $\text{CuO}_x$ /GC, (c) Pd/ $\text{CuO}_x$ /GC, and (d)  $\text{CuO}_x$ /Pd/GC electrodes in 0.3 M FA (pH = 3.5). Potential scan rate:  $100 \text{ mVs}^{-1}$ .

It is important to consider the inactivity of the unmodified GC electrode toward FAO [25, 43]. Inspection of Figure 5 reveals few interesting remarks:

- (i) The reaction pathway of FAO on the Pd/GC electrode (Figure 5(a)) proceeded exclusively via the dehydrogenation route in which  $\text{CO}_2$  was the only oxidation product with insignificant contribution of Pd poisoning by CO [44, 45]. This was evident from the appearance of a single oxidation peak around ca. 0 V, which assigned the direct oxidation of FA to  $\text{CO}_2$  [33].
- (ii) The electrodeposition of  $\text{CuO}_x$  directly at GC electrode surface did not show any activity for FAO (Figure 5(b)) to confirm the inertness of  $\text{CuO}_x$  as well as the GC substrate toward FAO.

- (iii) The deposition of PdNPs at GC electrode surface modified with  $\text{CuO}_x$  to obtain Pd/ $\text{CuO}_x$ /GC electrode (Figure 5(c)) a little bit shifts the current density of the oxidation peak ( $I_p$ ) to a higher value as compared to the Pd/GC electrode (Figure 5(a)). This was likely because it had a higher surface area of PdNPs (as discussed before in Figure 1).
- (iv) The electrodeposition of  $\text{CuO}_x$ NWs at the GC electrode surface previously modified with PdNPs (Pd/GC) to obtain  $\text{CuO}_x$ /Pd/GC electrode (Figure 5(d)) exhibited an enhanced performance toward FAO with an increase of  $I_p$  (ca. 1.5 fold) and a negative shift (ca. 30 mV) in the onset potential ( $E_{\text{onset}}$ ) of FAO. Recalling the smaller surface area of Pd in the  $\text{CuO}_x$ /Pd/GC electrode than in the Pd/GC electrode (compare Figures 1(a) and 1(d)), one may exclude the geometrical factor in the catalytic enhancement to suggest rather an electronic influence for  $\text{CuO}_x$ NWs that could presumably promote the charge transfer of the direct FAO at the electrode surface [17, 29, 46].

**3.2.2. Effect of Loading Level of  $\text{CuO}_x$ NWs at Pd/GC Electrode.** Optimization of the loading level of  $\text{CuO}_x$ NWs onto the Pd/GC electrode toward FAO has been investigated. A gradual decrease in the surface area of PdNPs reflected by a decrease in the intensity of the Pd oxide reduction and the  $H_{\text{ads/des}}$  peaks with the loading of  $\text{CuO}_x$ NWs have been observed. Table 1 summarizes the variation of the surface coverage ( $\Gamma$ ) of  $\text{CuO}_x$ NWs electrodeposited onto the Pd/GC electrode with time and its effect on the peak current of FAO,  $I_p$ . As obviously seen, the  $I_p$  increased systematically to a certain value with  $\Gamma$  and then decreased. This makes sense as the catalytic activity of the  $\text{CuO}_x$ /Pd/GC electrode is expected to increase with the loading level of  $\text{CuO}_x$ NWs as long as the accessible surface area of PdNPs is sufficient to sustain the adsorption of FA. However, beyond certain coverage (ca. 49%), the accessible surface area of PdNPs becomes limiting in estimating the overall rate of FAO.

**3.3. Stability of  $\text{CuO}_x$ /Pd/GC Electrode.** Stability of the Pd-based catalysts is another important parameter in the practical use of DFAFCs. Figure 6 shows current transients of the Pd/GC, Pd/ $\text{CuO}_x$ /GC, and  $\text{CuO}_x$ /Pd/GC electrodes in a 0.3 M aqueous solution of FA (pH = 3.5) at  $-0.1 \text{ V}$ . We did not include the data of the  $\text{CuO}_x$ /GC electrode as it did not show any activity toward FAO (Figures 5(b), 6(a), and 6(b)) (Pd/GC and Pd/ $\text{CuO}_x$ /GC electrodes, respectively) depicting a poor stability toward FAO, as can be obtained from the fast current drop with time, which agreed with other investigations in the literature [33, 46]. Interestingly, the  $\text{CuO}_x$ /Pd/GC electrode (Figure 6(c)) exhibited a higher current density in the activation polarization region, and its final current density was  $1.6 \text{ mA}\cdot\text{cm}^{-2}$  (compare it with  $0.07$  and  $0.3 \text{ mA}\cdot\text{cm}^{-2}$  obtained at the Pd/GC and Pd/ $\text{CuO}_x$ /GC electrodes). This may be assigned to the difference in the degree of poisoning of the three electrodes. The long-term

TABLE 1: Variation of the surface coverage ( $\Gamma$ ) of  $\text{CuO}_x$ NWs electrodeposited onto Pd/GC electrode with deposition time and its effect on  $I_p$  of FAO.

Dep. time (min)	Real surface area ( $S$ ) <sup>a</sup> ( $\text{cm}^2$ ) $\text{CuO}_x/\text{Pd}/\text{GC}$	( $\Gamma\%$ ) <sup>b</sup>	$I_p$ ( $\text{mA}\cdot\text{cm}^{-2}$ )
0	0.49	0	4.5
1	0.38	22	5.2
2	0.25	49	6.5
3	0.17	65	5.9
4	0.10	80	5.4

<sup>a</sup>As estimated from the charge consumed during the reduction of the surface oxide monolayer in 0.5 M  $\text{H}_2\text{SO}_4$  using a reported value of  $420 \mu\text{C}\cdot\text{cm}^{-2}$ . <sup>b</sup>The values of surface coverage ( $\Gamma = 1 - S_{\text{modified}}/S_{\text{unmodified}}$ ) were calculated for the various  $\text{CuO}_x/\text{Pd}/\text{GC}$  electrodes.  $S_{\text{modified}}$  and  $S_{\text{unmodified}}$  refer to the real surface area of the modified and the unmodified Pd/GC electrodes, respectively.

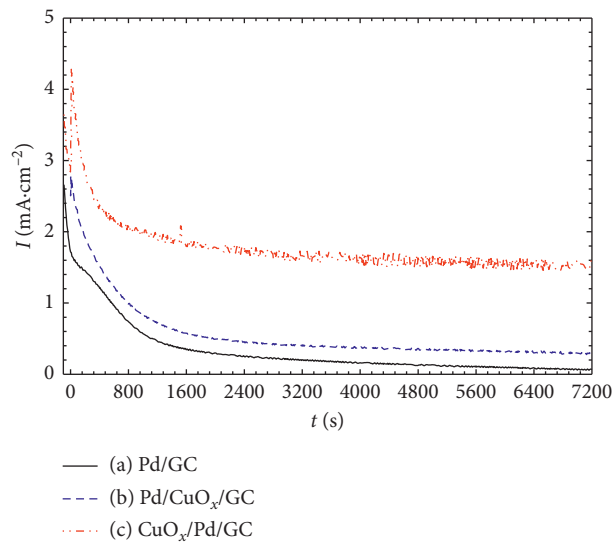


FIGURE 6: Current transients of (a) Pd/GC, (b) Pd/ $\text{CuO}_x$ /GC, and (c)  $\text{CuO}_x$ /Pd/GC electrodes in 0.3 M FA (pH = 3.5) at  $-0.1$  V.

poisoning rate ( $\delta$ ) could be calculated by measuring the linear decay of the current at intervals longer than  $500$  s using the following equation [47]:

$$\delta = \frac{100}{i_o} \times \left( \frac{di}{dt} \right)_{t > 500} \quad \% \cdot \text{s}^{-1}, \quad (1)$$

where  $(di/dt)_{t > 500}$  is the slope of the linear portion of the current decay and  $i_o$  is the current at the start of polarization back-extrapolated from the linear current decay, respectively. Fascinatingly, the  $\text{CuO}_x/\text{Pd}/\text{GC}$  electrode showed a lower poisoning rate (ca.  $0.002\% \cdot \text{s}^{-1}$ ) if compared to those obtained at both the Pd/GC (ca.  $0.007\% \cdot \text{s}^{-1}$ ) and Pd/ $\text{CuO}_x$ /GC electrodes (ca.  $0.005\% \cdot \text{s}^{-1}$ ). This also infers about how  $\text{CuO}_x$  promoted FAO at the Pd- $\text{CuO}_x$  modified catalysts.

Furthermore, some important electrocatalytic parameters, calculated from pseudosteady state current, such as time needed to oxidize a complete monolayer of FA and turnover frequency (TOF), have been estimated for the modified electrodes. Utilizing the charge associated with oxidation of a complete monolayer of FA, the complete oxidation of a FA monolayer at the Pd/GC and Pd/ $\text{CuO}_x$ /GC electrodes, respectively, needed 0.86 and 0.82 s on the basis of a two-site adsorption model, and the corresponding TOF was estimated as  $2.33$  and  $2.44 \text{ s}^{-1}$  [48], respectively. Interestingly, based on

the same adsorption model, the complete oxidation of a FA monolayer at the  $\text{CuO}_x/\text{Pd}/\text{GC}$  electrode needed 0.35 s, and the corresponding TOF was  $5.71 \text{ s}^{-1}$ . These parameters could infer about the superior catalytic activity of the as-prepared  $\text{CuO}_x/\text{Pd}/\text{GC}$  electrode. Table 2 summarizes the electrochemical parameters, extracted from Figures 5 and 6, of the  $\text{CuO}_x/\text{Pd}/\text{GC}$  compared with Pd/GC and Pd/ $\text{CuO}_x$ /GC electrodes. It also compares the data obtained in this investigation with the same data obtained from other previous investigations at which Pt surface was modified with nano- $\text{NiO}_x$  and IrNPs [49, 50].

Further, the inductively coupled plasma optical emission spectrometry (ICP-OES) has been utilized to specify if there was a loss in the catalysts' materials (Pd and Cu) after the prolonged electrolysis and to determine the amounts lost. A minor loss was obtained after 2 h of electrolysis in FA. Importantly, the loss of the active material toward FAO (Pd) in the Pd/GC electrode was ca. 1.5 of that obtained at the modified  $\text{CuO}_x/\text{Pd}/\text{GC}$  electrode confirming once again the superiority of the  $\text{CuO}_x/\text{Pd}/\text{GC}$  electrode toward FAO (Table 3).

#### 4. Conclusion

A novel binary catalyst composed of PdNPs and  $\text{CuO}_x$ NWs has been assembled onto the GC electrode for the efficient

TABLE 2: Electrochemical parameters, extracted from Figures 5 and 6, of the  $\text{CuO}_x/\text{Pd}/\text{GC}$  compared with  $\text{Pd}/\text{GC}$ ,  $\text{Pd}/\text{CuO}_x/\text{GC}$  electrodes in this study and other previously reported catalysts.

Electrode	$I_p$ ( $\text{mA}\cdot\text{cm}^{-2}$ )	$I_E = 6 \text{ mV}/\text{mA}\cdot\text{cm}^{-2}$	$E_{\text{onset}}$ (mV)	$\delta$ ( $\%\cdot\text{s}^{-1}$ )	TOF ( $\text{s}^{-1}$ )	References
$\text{Pd}/\text{GC}$	4.40	4.52	-230	0.007	2.33	This study
$\text{Pd}/\text{CuO}_x/\text{GC}$	5.00	4.93	-242	0.005	2.44	This study
$\text{CuO}_x/\text{Pd}/\text{GC}$	6.50	6.51	-260	0.002	5.71	This study
$\text{Pd}/\text{NiO}_x/\text{GC}$	5.20	5.29	-236	0.006	2.10	[49]
$\text{Pd}/\text{Ir}/\text{GC}$	3.30	3.25	-250	0.05	1.90	[50]

$I_E = 6 \text{ mV}$ : current density value at a same potential of 6 mV.

TABLE 3: Inductively coupled plasma optical emission spectrometry (ICP-OES) measured after 2 h of electrolysis in FA for  $\text{CuO}_x/\text{Pd}/\text{GC}$  and  $\text{Pd}/\text{GC}$  electrodes.

Analyte	Wavelength (nm)	Electrode	Conc (sample) ( $\text{mg}\cdot\text{L}^{-1}$ )
Pd	340.458	$\text{Pd}/\text{GC}$	0.032
		$\text{CuO}_x/\text{Pd}/\text{GC}$	0.021
Cu	324.752	$\text{CuO}_x/\text{Pd}/\text{GC}$	0.095

FAO. The deposition sequence of both PdNPs and  $\text{CuO}_x$ NWs onto the GC electrode influenced, to a great extent, the catalytic efficiency. The highest catalytic activity and stability were obtained at the  $\text{CuO}_x/\text{Pd}/\text{GC}$  electrode (in which  $\text{CuO}_x$ NWs were partially deposited,  $\Gamma \approx 49\%$ , onto the  $\text{Pd}/\text{GC}$  electrode). The enhancement in the electrocatalytic activity (increasing  $I_p$  of direct FAO by ca. 1.5-fold and decreasing  $E_{\text{onset}}$  by ca. 30 mV) and stability (decreasing both current decay and loss of active metal, by ca. 1.5-fold, during prolonged electrolysis) of the  $\text{CuO}_x/\text{Pd}/\text{GC}$  electrode was believed originating both from facilitating the direct oxidation of FA (increasing turnover frequency by ca. 2.5-fold) and minimizing the adsorption of poisoning species (by ca. 71.5%) at the electrode surface.

## Data Availability

The data used to support the findings of this study are available from the corresponding author upon request.

## Conflicts of Interest

The authors declare that there are no conflicts of interest regarding the publication of this paper.

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