

Research Article

Konjac Sponge Derived Carbon Flakes with Optimized Pore Structure for High-Performance Supercapacitor

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Lamellar activated carbons derived from Konjac sponges (KACs) have been successfully fabricated through a facile KOH activation method. By manipulating the activation temperature and KOH/C ratio, the achieved KACs exhibit ultrahigh specific surface area up to $\sim 3000 \text{ m}^2/\text{g}$ and hierarchical pore structure with tunable micro/mesopore distribution. Notably, KACs possess plenty of worm-shaped micropores formed by graphene stacking layers with the lateral distance close to size of hydrated electrolyte ions. Owing to optimized pore structure, high graphitization, and extra O/N doping, KACs exhibited much enhanced specific capacitance (253.0 F/g), superior rate ability (77% retention of capacitance at 10 A/g), and remarkable cycling stability (0.4% decay under 5 A/g after 2000 cycles) in the acid electrolyte. The mass production ability of KAC materials and the knowledge of correlation between texture properties and capacitive performance open new opportunities for the application of such novel biomass-derived carbons in supercapacitor devices.

1. Introduction

The shortage of fossil fuels and booming demand for sustainable energy call for the development of highly efficient energy storage solutions [1, 2]. Supercapacitors (SCs), as a complementary or independent charge storage system for electric vehicles or regenerated applications, have exhibited bright prospect due to the high power density, fast charging process, and excellent cycling performance [3]. Based on the surface adsorption/desorption of electrolyte ions, electrochemical double-layer capacitors (EDLCs) are one important component of commercial electrochemical capacitors owing to their low cost and high operation safeties [4]. Nevertheless, the relatively low-energy density of EDLCs has restricted their further application. Therefore, much efforts have been made to increase the energy storage capability of EDLCs in recent years [5, 6]. One of the most effective strategies is to improve the specific capacitance through sophisticated design of electrode materials. Carbon nano-materials are the dominant electrode materials for EDLCs. The past few decades have witnessed the development of various nanostructured carbon materials in SCs, ranging from the modified activated carbons (ACs) [7] to novel

carbon nanostructures based on carbon nanotube [8], graphene [9], or metal-organic framework- (MOF-) derived carbon [10].

ACs, by contrast, have been the mainstream electrode material in commercial SC devices due to the cost-effective and easy to scale up nature. To satisfy the demand of next generation SCs for both high-specific capacitance and rate ability, rational design of pore distribution/structure and optimization of surface modification have been urgently required for ACs while maintaining the considerable specific surface area (SSA) [11–13]. On one hand, micropores (pore size $< 2.0 \text{ nm}$) provide effective sites for ion adsorption-desorption, and mesopores ($2.0 \text{ nm} < \text{pore size} < 50 \text{ nm}$) play a dominant role in ion transportation [14]. Moreover, several studies have demonstrated the importance of sub-nanopores (pore size $< 1.0 \text{ nm}$) in largely increasing capacitance when the pore dimension matches the radius of the electrolyte ions [15, 16], which inspired attention on further optimizing the hierarchical pore structure. On the other hand, heteroatom (N, O, S, etc.) doping or group functionalization could produce surface redox reaction, therefore improving the accessibility by optimizing the wettability and introducing specific pseudocapacitance [13, 17, 18].

Furthermore, manipulation of the graphitization degree would increase the electrical conductivity, which facilitates the charge transfer especially at high rate [19, 20]. Therefore, development of well-designed ACs with high SSA, well-designed pore structure, heteroatom doping, and high graphitization degree opens new opportunities for this sort of traditional materials in future SC devices. Notably, biomass is an excellent precursor with the advantages of large amount, rich composition, and distinct primary texture [21]. Upon appropriate activation process, hierarchically porous ACs which inherit the original microstructure and heteroatom can be achieved [22, 23]. Thus, selection of appropriate carbon precursor and optimization of activation condition become the research hotspot in nowadays.

Konjac, which is composed of glucomannan with D-mannose and *i*-glucose in b(1,4) linkages, has been widely used as an agent for food ingredient in Asia [24]. Due to its low cost, easy for gelation, and good biocompatibility, Konjac glucomannan (KGM) has served as a gelling, texturizing, and film-forming agent in industries such as food, drug, and cosmetic [25]. Nevertheless, the application of ACs derived from Konjac-derived ACs in SCs has been rarely reported. Li et al. have demonstrated the first study of Konjac gel-derived porous biochar as electrode active materials in SCs, which exhibited the specific capacitance of 273.8 F/g at 1 A/g in 1 M KOH [26]. Additionally, “snow Konjac” is a sponge-like Konjac product consisting of KGM and rice flour (mass ratio of ~1 : 1) as main gradients. Taking advantages of the viscosity of KGM and swelling property of rice flour, such Konjac sponges process a loose and porous appearance with lamellar inner wall, which are expected to serve as a novel platform to achieve new structured ACs. Herein, we fabricated two-dimensional (2-D) graphitized carbon nanosheets from the Konjac sponge through the facile KOH activation method. By tuning the carbonization temperature and KOH/C ratio, Konjac sponge-derived ACs (KACs) with ultrahigh SSA (up to ~3000 m²/g at KOH/C > 3), tunable micro/mesopore volume, varied O/N doping, and high graphitization degree have been achieved. Impressively, such 2-D KAC nanosheets show hierarchical pore structure with coexistence of micro/mesopores. Especially, the micropores contain large amount of worm-shaped graphene stacking layers with the lateral distance close to the size of hydrated electrolyte ions, which greatly benefit the capacitive performance. Therefore, compared with non-activated samples, KACs exhibited much enhanced specific capacitance, superior rate ability, and remarkable cycling stability in acid electrolyte. The relationship between the capacitive performance (specific capacitance and rate ability) and the texture property (SSA, micropore volume, and mesopore volume) as well as the graphitization degree is also elucidated, which would largely benefit the design and practical application of such KAC materials in SCs.

2. Materials and Methods

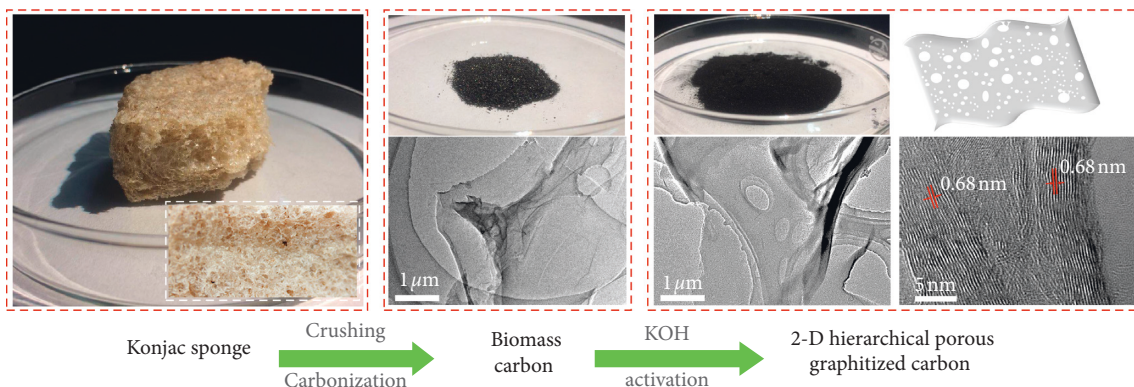
2.1. Chemicals. The Konjac sponge with the brand “Leshan” was bought from Jingdong Mall. Potassium hydroxide (KOH) and hydrochloric acid (HCl) used were obtained

from Aladdin (Shanghai, China). *N*-methylpyrrolidinone (NMP) was bought from Alfa Aesar.

2.2. Fabrication of Hierarchical Porous Carbon from Konjac Sponge. The Konjac sponge was firstly crushed and pre-carbonized in the tube furnace under N₂ atmosphere at 500°C for 2 h, and the achieved sample was mixed thoroughly with KOH (C/KOH ratio of 1 : 3). The mixture was then carbonized under N₂ flow at 700°C, 800°C, and 900°C for 2 h. The resultant products were marked as Konjac derived activated carbon (KAC)- 700, 800, and 900, respectively. The precarbonized sample carbonized under 700°C without the addition of KOH was denoted as UAC. To evaluate the effect of carbonized temperature on the capacitive performance, the precarbonized carbon was mixed with different C/KOH ratios (1 : 1, 1 : 2, and 1 : 4) before the further carbonization at 800°C under N₂ atmosphere. The derived samples were denoted as KAC-1, 2, and 4, respectively. KAC-800 was equal to KAC-3. All the KAC samples were washed with 2 M HCl to remove the residual KOH and centrifuged with deionized water until the solution reached neutral condition. After drying in vacuum, the KAC materials were obtained. The synthetic process is shown in Scheme 1.

2.3. Characterization. The morphology and high-resolution imaging were obtained from a Tecnai G2 F20 U-TWIN. Material phase analysis was conducted using an X-ray diffractometer (D/max-TTRIII). Information of chemical and elemental composition was achieved from X-ray photoelectron spectroscopy (XPS). Raman analysis was performed on a Renishaw inVia plus system at 513 nm laser excitation. Texture and pore information were recognized based on N₂ adsorption-desorption isotherms from a Micromeritics ASAP 2010 analyzer.

2.4. Electrochemical Testing. The potential of KAC-based samples as SC electrode materials was evaluated on a CHI660E workstation. The related measurements were performed through a three-electrode system with an electrolyte of 1 M H₂SO₄. KAC-based samples (KAC-700, 800, and 900 and KAC-1 and 2) were ground with specific amount of acetylene black and polyvinylidene fluoride (mass percentage is 8 : 1 : 1) in *N*-methyl pyrrolidone (NMP) to form slurry, which was dispensed on graphite flake to form an area of 1 cm × 1 cm. The electrodes were dried in vacuum (80°C, 12 h). The mass loading was obtained by measuring the mass of the blank substrate and the dried working electrode. The Pt plate counterelectrode, as a prepared working electrode, and Hg/Hg₂SO₄ reference electrode formed the three-electrode system. Cyclic voltammetry (CV) and galvanostatic charge-discharge (GCD) results were achieved in the potential scope of -0.6~0.3 V. Electrochemical impedance spectra (EIS) were achieved under the condition of 5 mV and 100 kHz~0.01 Hz. The specific capacitance (Cs, F/g) was calculated based on the GCD curve according to the following equation:



SCHEME 1: The synthetic procedures for hierarchical porous KAC materials.

$$C_s = \frac{It}{\Delta V m}, \quad (1)$$

where I and t are the charge-discharge current and discharging time, respectively, ΔV is the potential range, and m is the mass of the active material.

3. Results and Discussion

The pristine Konjac sponge precursor possessed laminar network structure in macroscopic view. TEM imaging provided intuitional information about the morphology and porous structure of the samples. As shown in Figure S1, the UAC sample exhibited very thin lamellar structure upon crushing and carbonization. The enlarged HR-TEM images indicated the amorphous nature of the nonactivated carbon flakes, which contained nanosized pores. In contrast, KOH activation gave rise to tremendous modification towards the KACs both in microstructure and graphitization. Interestingly, worm-shaped graphene bunches and amorphous area coexisted in KAC-based samples (Figures 1 and S2). The distance between graphene layers was 0.67 and 0.68 nm in KAC-700 and KAC-900, respectively (Figure S3). KAC-800 was dominated by the distorted short-range ordered crystalline with a lateral distance of ~ 0.67 and ~ 0.79 nm. As shown in Figure S4, staking graphene layers with larger lateral distance were clearly observed in KAC-1 (0.79 nm), KAC-2 (0.85 nm), and KAC-4 (0.75 nm). The enlargement of interlayer spacing for KAC-700 was further confirmed from selected-area electron diffraction (SAED) pattern (Figure S5(a)). Compared with the amorphous UAC carbon, the graphene crystallization zone indicated the increased graphitization of the KAC samples, which would largely benefit the capacitive performance arising from better electrical conductivity. Notably, the KOH etching effect became particularly obvious at a high temperature of 900°C and a high KOH/C ratio of 4:1, with the observation of the rough surface and plentiful mesopores. The negligible observation of the change of carbon lattice distances for KAC-4 may be due to the less-ordered crystalline area from the existence of great deal of mesopores (Figure S5(b)). Therefore, the hierarchical porous structure consisting of large amount of graphene layer subnanopore and a portion of mesopores formed due to activation process.

Furthermore, the XRD pattern of UAC showed broad peaks at $2\theta = 23^\circ$ and 43° (Figure 2(a)), corresponding to the diffraction signal of (002) and (101) plane of amorphous carbon [27, 28]. KAC-based samples, in comparison, exhibited a negligible peak at 23° and largely reduced diffraction at 43° (Figures 2(a) and 2(b)), suggesting the rearrangement of the original amorphous carbon and introduction of large amount of pores. The steep increase in low-angle scatter in KAC samples compared with UAC originates from the enriched micro/mesopores [29]. As is shown in HRTEM, the graphene interlayer distance was much larger than the graphite (002) d-spacing of 0.34 nm, which may arise from the separation effect due to K insertion/intercalation into graphene layers [30]. The Raman spectra quantitatively demonstrate the graphitization degree of carbon materials (Figures 2(c) and 2(d)). KAC samples showed two Raman peaks at 1357 cm^{-1} and 1594 cm^{-1} characteristic of polycrystalline and crystal graphite, respectively [31, 32]. The intensity ratio of D-band and G-band (I_D/I_G) for KAC-700, 800, and 900 is 0.96, 0.96, and 1.00, respectively, while that for KAC-1, 2, and 4 were 0.95, 0.99, and 0.74. The much lower I_D/I_G value of KAC samples compared with commercial activated carbon (~ 1.92) demonstrated the high graphitization degree [33]. This result also illustrated that the disorder degree of carbon increased with the carbonization temperature. The sharper peak with temperature and KOH/C was consistent with the larger graphene separation observed from HRTEM images. On one hand, the high content of defective carbon provided more active sites or active surface areas, giving rise to improved capacitance. On the other hand, graphitized carbon increases the electric conductivity which benefited the rate capability [20]. Therefore, the optimized capacitive performance is called for the good intercoordination of both effects.

XPS spectra further provided the chemical composition and element binding information. As shown in Figure 3, the high-resolution C1s spectra of KAC-based samples consisted of a peak centered at 284.7 eV with small tail at the high-binding energy side and a shoulder peak in the range of 288.0~292.8 eV. The former peak can be deconvoluted into sp^2 -bonded carbon (284.7 eV) with a small tail containing sp^3 -bonded carbon (285.8 eV). The shoulder peak consisted of three individual component peaks, corresponding to

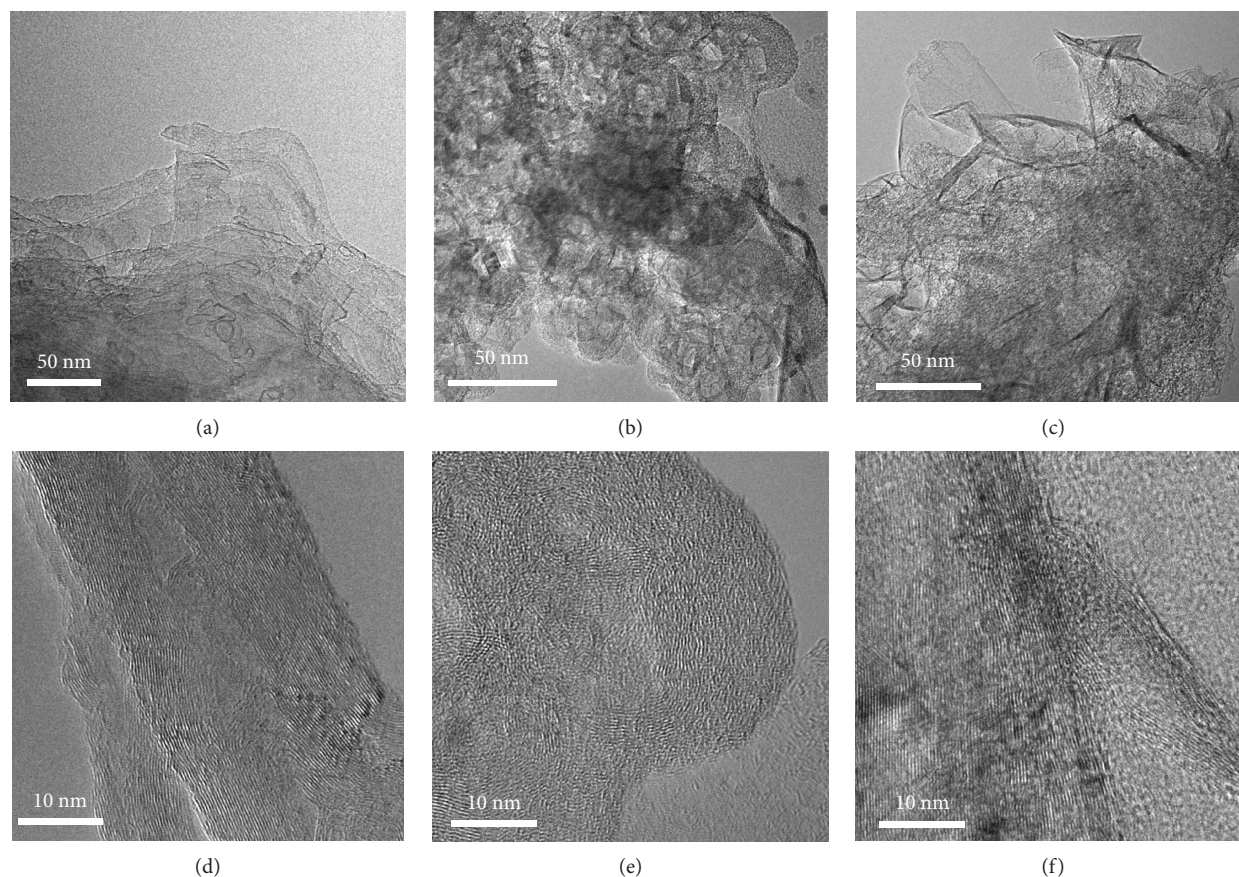


FIGURE 1: Transmission electron microscopy (TEM) images of (a, d) KAC-700, (b, e) KAC-800, and (c, f) KAC-900.

carbonyl group C=O (287.0 eV), carboxyl O-C=O (289.2 eV), and a satellite peak of $\pi-\pi^*$ (291.4 eV), respectively [34, 35]. The overwhelmingly strong peak of sp^2 carbon demonstrated the formation of graphene-like carbon sheets with high graphitization, which agreed well with the Raman results. As shown in Figure S6, O1s spectra of KAC-based samples indicated the presence of C=O quinone-type groups (O-I), C-OH phenol groups and/or C-O-C ether groups (O-II), and chemisorbed oxygen (COOH carboxylic groups) and/or water (O-III), respectively [36]. Table S1 shows that KACs were composed of C and O and trace amount of N. With the activation temperature and the KOH/C ratio increasing, the content of carbon increased and the oxygen decreased (Table S1). The decreased total intensity of oxygenated groups is also observed from Figure S6. The existence of carbonyl/carboxyl groups and small amount of N benefited the improvement of the wettability, therefore facilitating the accessibility of the electrolyte ions. Also, the oxygenated groups provided pseudocapacitance for the EDLC-dominated capacitance of KAC-based samples [37].

As is revealed from TEM images, KOH etching process produced abundant hierarchical pores in the Konjac-derived carbon substrate. Therefore, BET characterization was conducted to further evaluate the texture properties and the effect of KOH amount and activation temperature on the pore structure of KAC materials. As shown in Figures 4(a) and 4(c), UAC and all the KAC samples exhibited steep N_2

adsorption at very low relative pressure (p/p_0), which was the characteristic of the type-I isotherm [38]. This result indicated the existence of large amount of micropores in these samples. Notably, KAC-900 and KAC-4 possessed a H_4 -type hysteresis loop at $p/p_0 > 0.4$, suggesting the presence of mesopores [11]. The pore size distribution curves (Figure 4(b)) showed that all KAC-700, 800, and 900 samples possessed subnanopores centered at ~ 0.68 and 0.80 nm, respectively. This was consistent with the lateral distance achieved from HRTEM images. Along with the subnanopores, large micropores of ~ 1.20 nm and large amount of mesopores coexisted in these samples. KAC-900 exhibited larger mesopores (~ 2.60 nm) than KAC-700 and 800 (~ 2.0 nm). In contrast, UAC only showed a pore peak centered at ~ 1.17 nm with very low pore volume, indicating the introduction of both subnanopores and mesopores through the KOH activation process. Moreover, changing the KOH/C could also modulate the pore structure of KAC materials. KAC-1 and 2 showed similar pore distribution with KAC-800, while KAC-4 exhibited obvious introduction of enlarged mesopores of 2.76 and 3.42 nm (Figure 4(d)). Previous studies evidenced that subnanopores with the size matching up with that of electrolyte ions could greatly enhance the specific capacitance [15, 39, 40]. Considering the reported size of hydrated SO_4^{2-} ion (0.533 nm) [41, 42], KAC-700, KAC-2, and KAC-3 containing large fraction of subnanopore sized at ~ 0.68 nm were expected to have highly

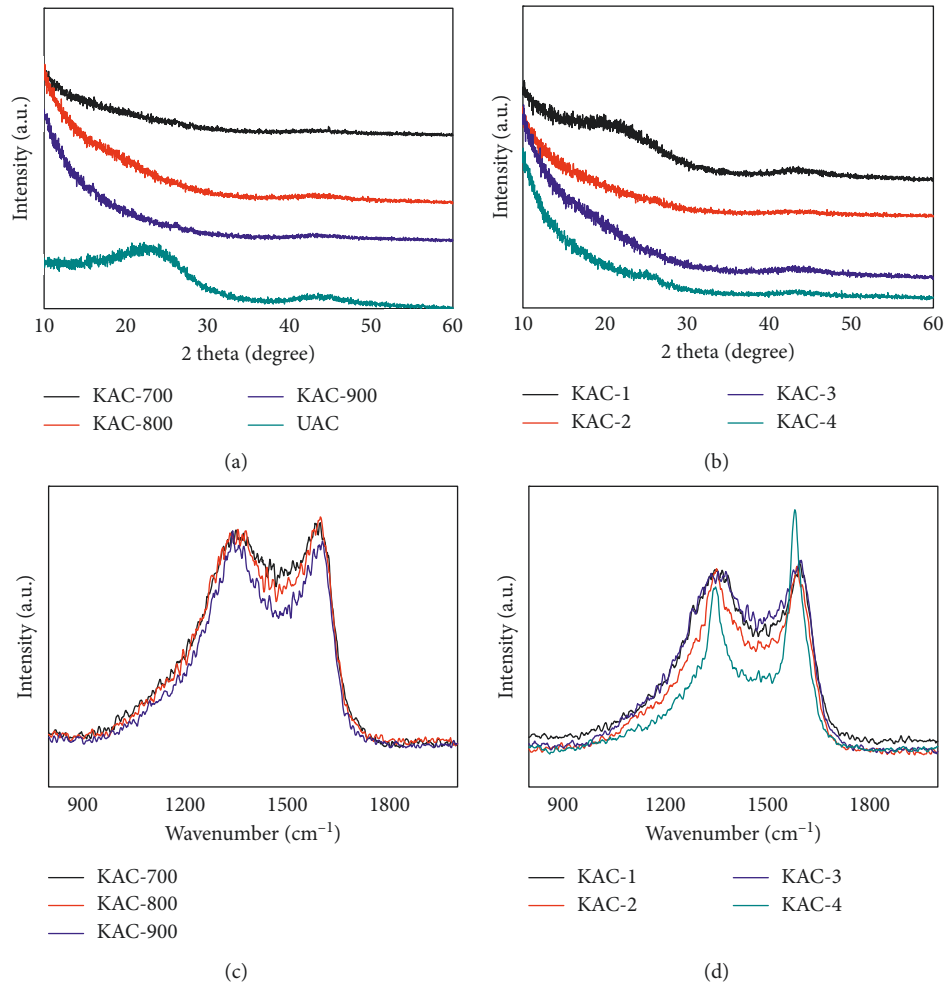


FIGURE 2: XRD patterns of KAC materials activated under (a) different temperatures and (b) different KOH/C ratios in comparison with UAC, (c) and (d) Raman spectra of KAC-based materials.

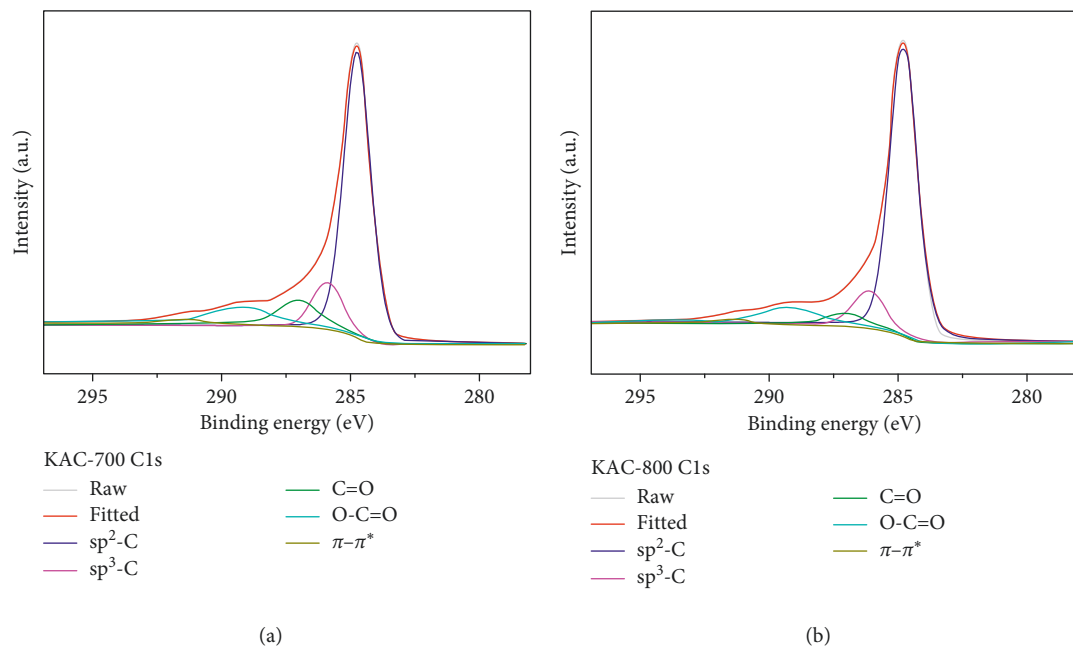


FIGURE 3: Continued.

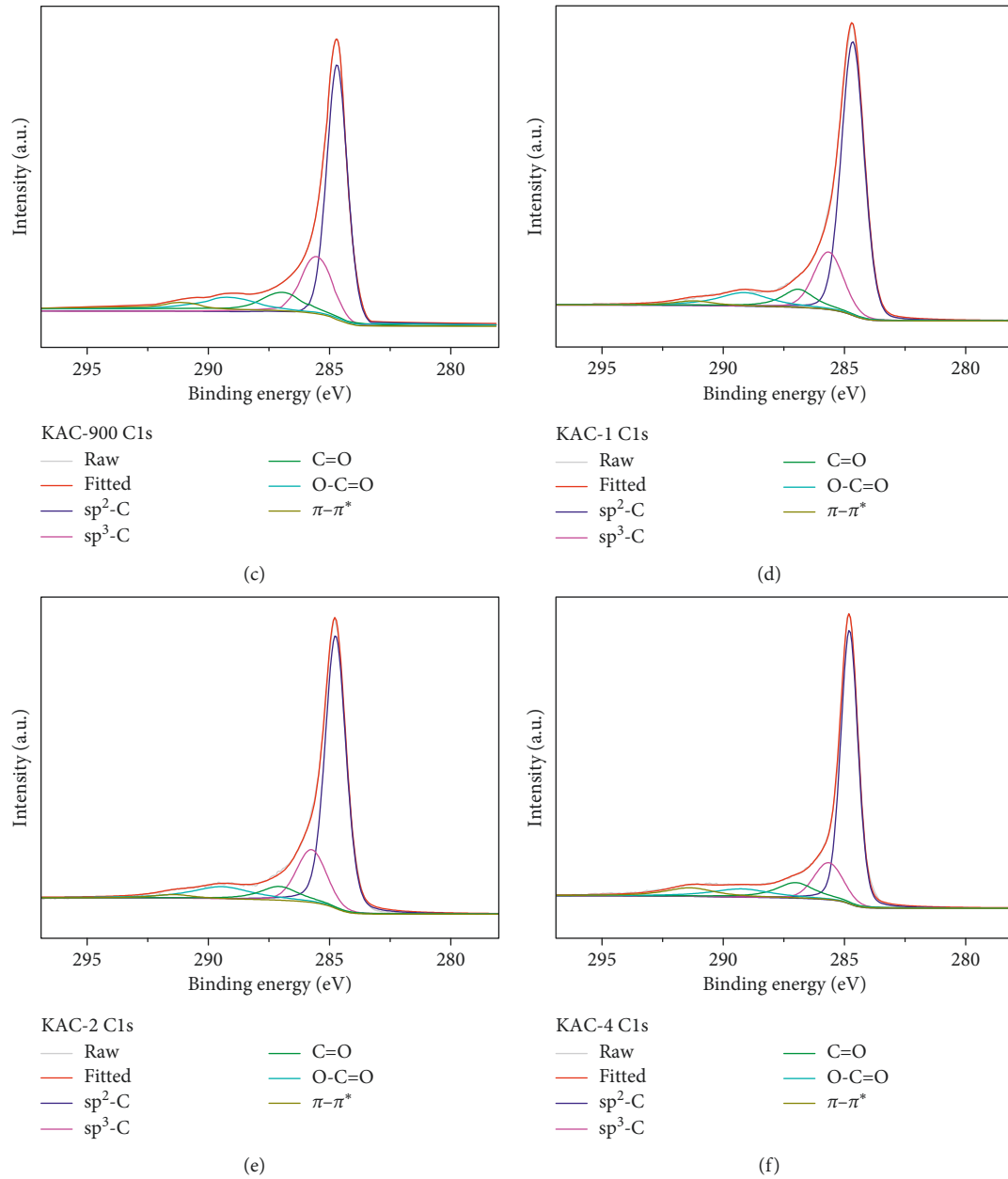


FIGURE 3: XPS C1s spectra of KAC-based samples.

effective surface area for ion adsorption-desorption, yielding improved specific capacitance. The worm-shaped pores formed from the carbon layers also contributed to the enhanced energy storage capability [43]. Despite the absence of macropores, the thin lamellar KAC materials would shorten the ion transportation route which reduced the diffusion impedance. Detailed specific surface area (SSA) and pore volume statistics are summarized in Table 1. At $\text{KOH}/\text{C} = 3$, KAC materials possessed ultrahigh SSA up to $\sim 3117 \text{ m}^2/\text{g}$ (KAC-700 and 900), which was about ~ 21 times and ~ 1.9 times higher than UAC and commercial-activated carbon, respectively. This result evidenced the efficiency of etching reaction between KOH and C precursor ($6\text{KOH} + \text{C} \leftrightarrow 2\text{K} + 3\text{H}_2 + 2\text{K}_2\text{CO}_3$) [29, 44]. As the ratio increases, SSA based on the BET model (S_{BET}) increased. With KOH/C ratio larger

than 3 : 1, SSA did not change much with the increase of the ratio or the carbonization temperature. But the pore structure including the distribution and relevant pore volume of micropores and mesopores exhibited large differences (Table 1). According to the density functional theory (DFT) method, KAC samples achieved at lower carbonization temperature (700 and 800°C) and low KOH/C ratio (1 : 1 and 2 : 1) showed SSA with 90%~93% contribution from micropores. Nevertheless, the portion of micropores decreased to 71% and 56% under the formation of mesopores for KAC-900 and KAC-4, respectively. The pore volume was also an important parameter to evaluate the capacitance and rate ability of EDLC-type SCs. Compared with the UAC sample, the KOH activation process under different conditions gave rise to the KAC materials with

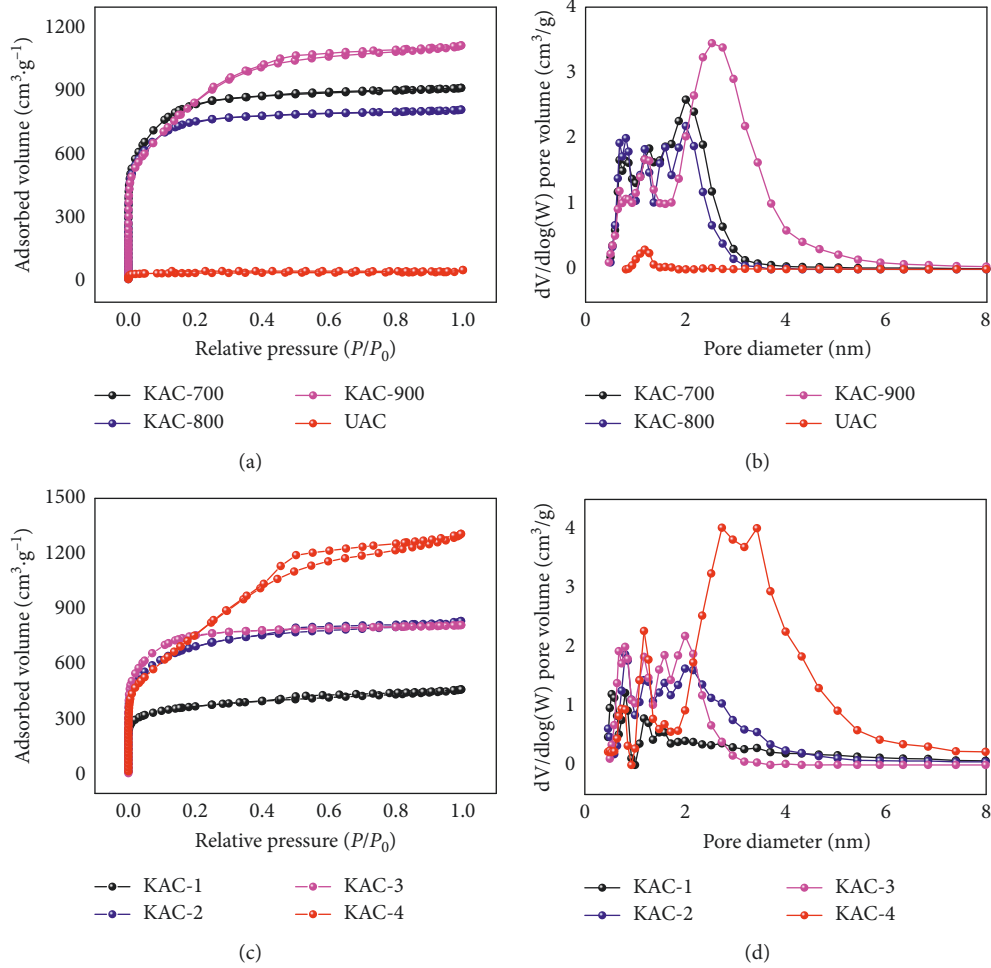


FIGURE 4: Nitrogen adsorption/desorption isotherms (a, c) and DFT pore size distribution curve (b, d) of KAC-based samples in comparison with UAC.

TABLE 1: Texture properties of the KAC samples achieved under different conditions and UAC.^a

Sample	S_{BET} (m ² /g)	S_{DFT} (m ² /g)	S_{DFT} (m ² /g)			Pore vol (cm ³ /g)		
			$d < 2$ nm	$d < 2$ nm	SSA_{mic} (%)	$d < 2$ nm	$d > 2$ nm	$V_{\text{mes}}/V_{\text{mic}}$
KAC-700	3117	1952	1757	195	90	0.94	0.24	0.25
KAC-800	2841	1861	1731	126	93	0.88	0.16	0.18
KAC-900	3116	1895	1352	543	71	0.70	0.77	1.10
KAC-1	1382	1178	1094	84	93	0.43	0.16	0.37
KAC-2	2557	1612	1402	210	87	0.76	0.31	0.41
KAC-4	2888	1687	949	738	56	0.49	1.24	2.53
UAC	152	70.5	68.3	2.2	97	0.05	0.005	0.10

^a S_{BET} and S_{DFT} : the specific surface area (SSA) based on Brunauer–Emmett–Teller (BET) and density functional theory (DFT) method; SSA_{mic} : the microporous surface area; $V_{\text{mes}}/V_{\text{mic}}$: the ratio between mesopore and micropore volume.

highly tunable pore structure. The KAC samples had very high pore volume compared with the biomass derived carbon counterparts [34, 45, 46]. Notably, KAC-700 possessed the highest micropore volume of 0.94 cm³/g while KAC-4 had the highest ratio between mesopore and micropore volume up to 2.56. The large microporous SSA and pore volume would benefit the charge storage capability due to high electroadsorption interaction between electrolyte ions and accessible pore surface. The high ratio of $V_{\text{mes}}/V_{\text{mic}}$

guaranteed the high-rate transportation of ions from solution to the active surface with the large fraction of mesopores acting as ion channels [47]. Therefore, the above results indicated the advantage of the Konjac sponge as the precursor for activated biomass carbon materials, and the efficiency of KOH etching treatment in providing optimized SC electrode materials with modulated SSA and pore structure. The series of KAC samples not only provided excellent model materials for studying the dependence of

supercapacitor performance on the structural properties of carbon materials but also paved the way for desirable carbon materials for high-performance SCs.

The capacitive performance of KACs as electrode materials for SC application was tested in the aqueous solution of 1 M H_2SO_4 . As shown in Figures 5(a), 6(a), and S7, all KAC samples exhibited rectangular-shaped CV curves at 5 mV/s and showed little deviation with the scan speed up to 100 mV/s, indicating the ideal EDLC behavior of the electrodes. The symmetric and triangular charge-discharge curves provided further confirmation for the good rate performance (Figures 5(b), 6(b), and S7). The small humps in CV curves originated from the pseudocapacitance due to O/N-heteroatom doping. It can be obviously noticed that KOH-activated carbons showed tremendously enhanced specific capacitance in comparison with UAC. The exact capacitance value was determined from GCD data (Table S2). When tuning the activation temperature, KAC-700 possessed the largest capacitance (235.1 F/g) at 0.5 A/g. Upon tuning the KOH/C ratio, KAC-2 and KAC-3 showed larger capacitance of 227.0 and 223.0 F/g, respectively. In overall, the modulation of KOH activation condition towards capacitance can be understood based on the previous analysis of pore information. KAC-700 had the largest microporous SSA (1754 m^2/g) and the maximum micropore volume (0.94 cm^3/g), followed by KAC-800 ($S_{\text{mic}} = 1731 \text{ m}^2/\text{g}$ and $V_{\text{mic}} = 0.88 \text{ cm}^3/\text{g}$ based on DFT calculation) and KAC-2 ($S_{\text{mic}} = 1402 \text{ m}^2/\text{g}$ and $V_{\text{mic}} = 0.76 \text{ cm}^3/\text{g}$). The large microporous surface provided abundant active sites for trapping the electrolyte ions, which greatly benefited the energy storage. Moreover, the pore structure played an important role in improving the specific capacitance. On one hand, KAC-700, KAC-2, and KAC-3 contained large fraction of the subnanopore sized at 0.68 nm. The well-fitted dimension with the electrolyte ion would benefit the energy storage [41, 42]. On the other hand, the moderate I_D/I_G demonstrated the existence of the defective carbon, which increased the exposed active surface for better charge storage performance.

Furthermore, rate capability was also a very important property to evaluate the capacitive performance of materials. The dependence of specific capacitance on the current density is shown in Figures 5(c) and 6(c). KAC-900 and KAC-4 showed remarkable capacitance retention of $\sim 77\%$ with the current density increasing from 0.5 A/g to 10 A/g, which was better than that of the rest KAC samples (70%–72%). The GCD curves of the KAC samples under different current densities are shown in Figure S8. For UAC, 71.4% of the original capacitance decayed at 10 A/g (Figure S9). The previous study revealed that the rate capability was dominated by the mesopore/micropore volume ratio [11, 47]. As shown in Figure S10, the trend of rate performance for KAC-based materials was in overall consistent with that of the $V_{\text{mes}}/V_{\text{mic}}$ values. The improved amount of mesopores in KAC-900 and KAC-4 served as ion transportation channels from the solution to the active surface of micropores, which guaranteed the smooth charge transfer at high-charge density [48]. Besides, the overall good rate performance of other KAC samples without much mesopores can be explained by the existence of large micropores sized at ~ 1.20 and ~ 1.59 nm, which were reported to facilitate the ion

diffusion process such as mesopores [27, 49]. The high graphitization degree gave rise to good conductivity of KAC materials, therefore leading to fast electron transport with low resistance. Furthermore, the EIS measurement provided deeper insight into the impedance referring to the ion diffusion and charge transfer. The Nyquist plots of KAC samples (Figures 5(d) and 6(d)) consisted of a semicircle at the high frequency with the real-axis intercept, followed by the Warburg resistance in the mid-to-high frequency region and a steep line in the low-frequency region. The small value of the intercept indicated the low equivalent series resistance (ESR), which was in good agreement with the negligible IR drop in the discharging curves of all the KAC samples [15, 50]. The semicircle diameter corresponded to the charge transfer resistance (R_{ct}) [46]. KAC-900 and KAC-4 showed very low R_{ct} value of ~ 0.11 and ~ 0.30 Ohm, respectively. Moreover, the shorter projected length of the 45° Warburg diffusion line of KAC-900 and KAC-4 confirmed the reduced diffusion resistance of electrolyte ions owing to the mesopores and large micropores with high pore volume and graphene layers' pathways with large layer distance [37]. The above results confirmed the high capacitance retention of KAC-900 and KAC-4 at high rate. Notably, the extended spike in the low frequency was typical of capacitive behavior. The closer the straight line to 90° , the better the capacitive performance [51, 52]. By combining pore structure characterization and EIS test, the tendency of specific capacitance and rate capability of KAC samples can be well understood. Our results were compared with previous studies, and the detailed literature review of the specific capacitance is shown in Table S3.

To evaluate the cycling stability of such Konjac-derived carbon materials, charge-discharge performance was tested on KAC-700 under 5 A/g for 2000 cycles. Figure 7 demonstrates a very good cycling performance with 99.6% retention of the initial capacitance. The slight increased capacitance arose from the activation of surface in the early stage of the charge-discharge process. The inset images showed the stable charge/discharge curves from the very beginning to the last few cycles. The superior long-range cycling performance confirmed that KAC-based samples were excellent electrode materials for the applications of high-efficient and stable supercapacitors.

Above all, KAC-700 exhibited the highest specific capacitance while KAC-900 and KAC-4 showed the best rate capability among the series of KAC materials. Combining structure and composition analysis, the dependence of the capacitive performance on the texture properties and chemical composition can be revealed as follows. For KAC-700, the highest microporous volume (0.94 cm^3/g) provided plentiful active sites for ion storage. The optimized subnanopore (~ 0.68 nm) further boosts the capacitance. The relatively low carbonization temperature maintained large amount of O and trace N, giving rise to additive pseudocapacitance. For KAC-900 and KAC-4, high activation temperature and KOH dose introduced large amount of mesopores of 2.76 and 3.42 nm with narrow distribution, which act as the ion transfer channels and reduce the diffusion resistance. The volume ratio between mesopores and

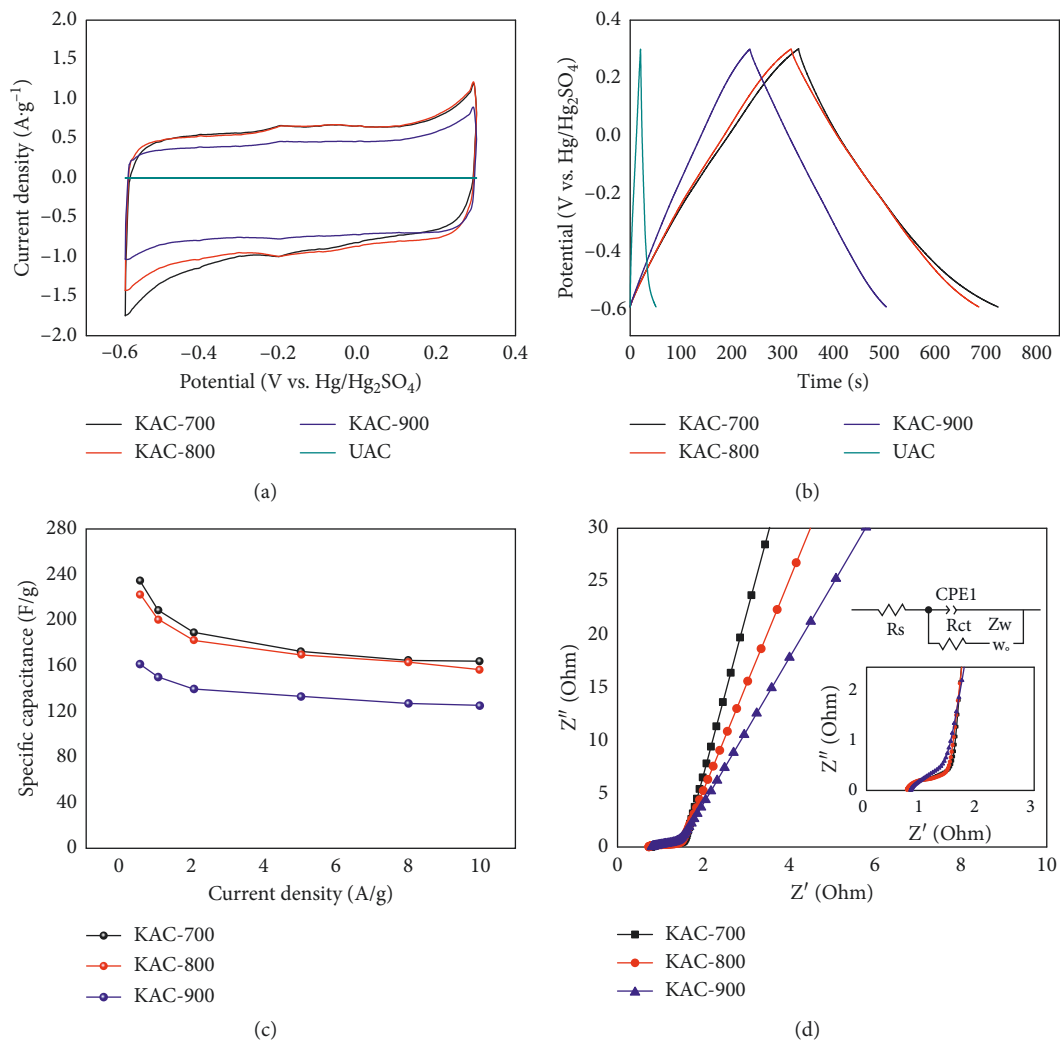


FIGURE 5: Electrochemical performance of KAC-700, 800, and 900 in 1 M H₂SO₄ aqueous solution. (a) CV curves at the scan rate of 5 mV/s, (b) GCD curves at 0.5 A/g, (c) specific capacitances under different current densities calculated from GCD data, and (d) Nyquist plots of the achieved samples, the inset images show enlarged impedance at the high frequency and the equivalent circuit.

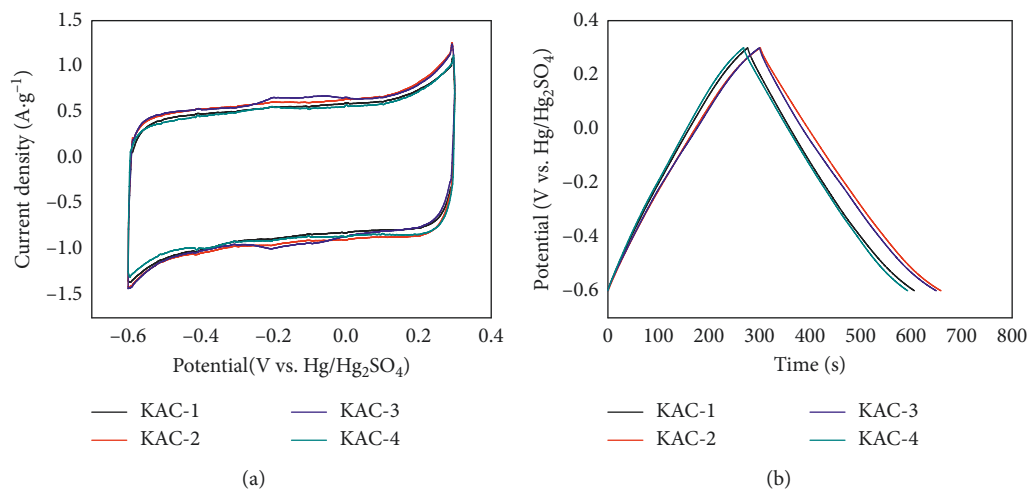


FIGURE 6: Continued.

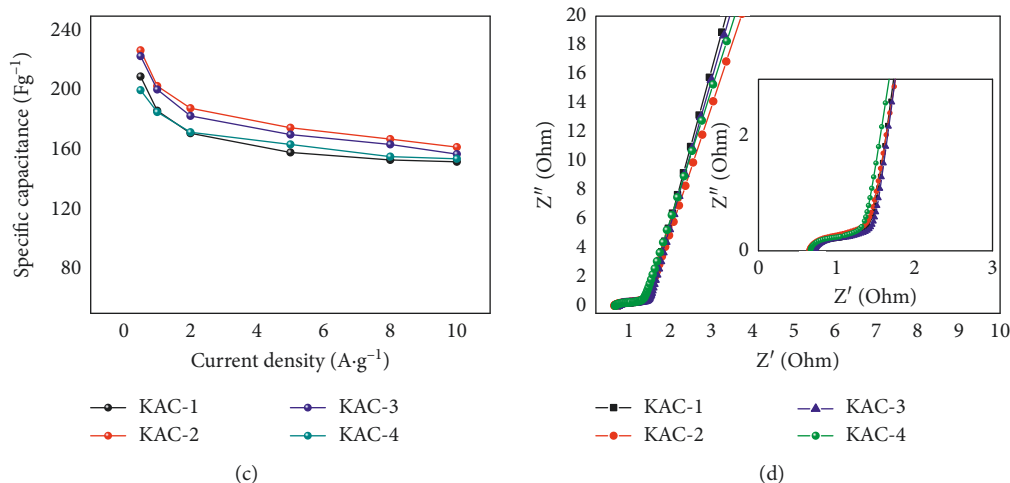


FIGURE 6: Electrochemical performance of KAC-1, 2, 3, and 4 in 1 M H₂SO₄ aqueous solution. (a) CV curves at the scan rate of 5 mV/s, (b) GCD curves at 0.5 A/g, (c) specific capacitances under different current densities calculated from GCD data, and (d) Nyquist plots of the achieved samples, the inset images showed enlarged impedance at the high frequency.

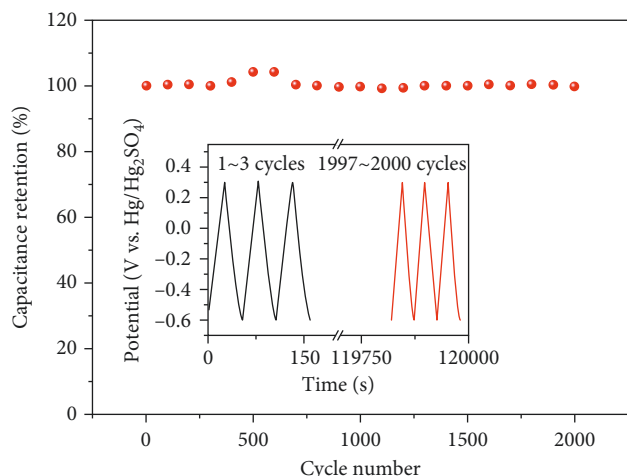


FIGURE 7: Cycling performance of KAC-700 under 5 A/g for 2000 cycles. The inset image shows the GCD curves in the first and last 3 cycles.

micropores is elevated compared with other KAC counterparts, indicating the better rate ability. The decreased O content and improved electrical conductivity benefit the charge transport especially at the high rate.

4. Conclusions

In summary, we have prepared the hierarchically porous ACs from the Konjac sponge precursor. Upon KOH activation, the KAC samples inherit the lamellar morphology of the raw materials and possess coexisted micro/mesopores with narrow size distribution. Remarkably, graphene-like layers are observed in KAC-based samples, with the stacking distance comparable to the size of the electrolyte ion. KAC-700 with rich O doping showed the highest specific capacitance owing to the largest ion accessible microporous volume with the optimized pore size. KAC-900 and KAC-4

possessed superior rate capability with the plentiful mesopores. Taking KAC-700 as an example, excellent cycling stability was obtained with only 0.4% decay of original capacitance under the current density of 5 A/g. The abundant source and simple synthetic procedure guarantee the scale-up production of such KAC materials, which would become novel and promising electrode materials for SC devices. The results achieved in this study can serve as guidance for the design and development of ACs with desirable porous structure, graphitized microstructure, and high SC performance.

Data Availability

The data used to support the findings of this study are available from the corresponding author upon request.

Conflicts of Interest

The authors declare that they have no conflicts of interest.

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Supplementary Materials

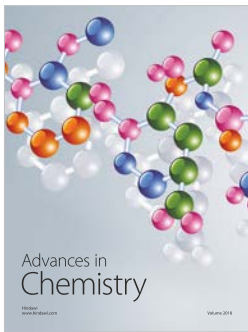
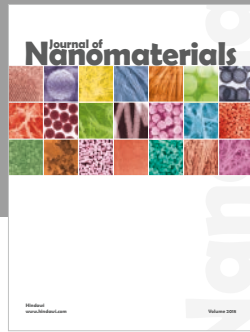
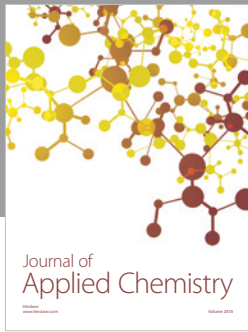
Figure S1. TEM images of UAC samples under different magnifications. Figure S2. Transmission electron microscopy (TEM) images of (a and d) KAC-1, (b and e) KAC-2 and (c and f) KAC-4. Figure S3. Transmission electron microscopy (TEM) images of (a) KAC-700, (b) KAC-800 and (c) KAC-900, with the related digital micrograph showing the lateral distance information, the numbers demonstrated the lateral distance in the selected area, scale bar = 10 nm. Figure S4. TEM images of (a) KAC-1,

(b) KAC-2 and (c) KAC-4, with the related digital micrograph showing the lateral distance information, the numbers demonstrated the lateral distance in the selected area, scale bar = 10 nm. Fig S5. Selected-area electron diffraction pattern of (a) KAC-700 and (b) KAC-4. Figure S6. O1s spectra of KAC based samples. Figure S7. CV curves of KAC based materials under scan rates of 1 mV/s, 5 mV/s, 10 mV/s, 20 mV/s, 50 mV/s, 100 mV/s. Figure S8. GCD curves of KAC based materials under the current densities of 0.5 A/g, 1 A/g, 2 A/g, 5 A/g, 8 A/g and 10 A/g. Figure S9. GCD curves (a) and the calculated specific capacitance (b) of UAC sample under different current densities. Figure S10. The microporous specific surface area (SSAmic) based on DFT model (Left Y axis), the microporous volume and ratio between mesoporous and microporous volume (V_{mes}/V_{mic}) (Right Y axis) of KAC based samples. Table S1. Elemental analysis results (atomic%) based XPS data. Table S2. The specific capacitance calculated from the discharging curves of KAC based samples and the capacitance retention under 10 A/g compared with that of 0.5 A/g. Table S3. Specific capacitance of KAC based materials using three-electrode cells reported in literatures. (Supplementary Materials)

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